

➤ *Effect of Solvent Polarity*

Just like the case of aliphatic nucleophilic substitution reactions, the raise in solvent polarity boosts the chances of the SE_1 pathway by supporting the ionization because of the better solvation of carbanions. However, if SE_2 and SE_1 reactions are competing with each other in parallel propagation, then less polar solvents favor the SE_2 pathway and polar solvents favor the SE_1 mechanism. Finally, If the nucleophilic character of the solvent is very small, the electrophile with properly placed assisting functionality might support the reaction; and therefore, motivating the reaction towards the SE_i pathway; otherwise, solvent polarity has little to no effect upon SN_i reactions.



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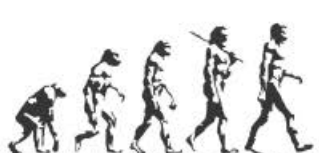
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A TEXTBOOK OF ORGANIC CHEMISTRY

Volume I

MANDEEP DALAL



First Edition

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